



Original Article

First Principles Study of the Electron Transport Properties of Buthane-dithiol Nano-Molecular Wire

Roya. Ahmadi ¹ , Masoud. Darvish Ganji ²

¹ Department of Chemistry, Faculty of Basic Sciences, Islamic Azad University, Yadegar Emam Khomeini Branch, Tehran, Iran

² Department of chemistry, Azad University of Ghaemshahr, Iran

Received: 5 March 2014; Accepted: 20 April 2014

Abstract:

We report a first-principles study of electrical transport in a single molecular conductor consisting of a buthane-dithiol sandwiched between two Au (100) electrodes. We show that the current was increased by increasing of the external voltage biases. The projected density of states (PDOS) and transmission coefficients (T(E)) under various external voltage biases are analyzed, and it suggests that the variation of the coupling between the molecule and the electrodes with external bias leads to the increase of the current. Therefore, we propose that the most origin of electron transport mechanism in molecular devices is caused by the characteristics of both the molecule and the electrodes as well as their cooperation, not necessarily only by the inherent properties of certain species of molecules themselves.

Keywords: single molecular conductor, Au(100) electrodes, DFT, buthane-dithiol

(*) Corresponding Author: e-mail: roya.ahmadi.chem@hotmail.com

Submit the manuscript to <http://sheminovin.ir>

Introduction:

One of the most important goals in the engineering of electronic devices is the miniaturization. The present technologies of microstructuring semiconductor material, especially silicon, are expected to reach their limits in the next decade. The smallest transistors are of about 100 nanometers wide, and tens of millions of them may be integrated on a single chip. Below this size the controlled doping becomes more and more difficult. The next important step in the miniaturization might be done by reducing the electronic devices to the scale of molecules. This new field of physics and chemistry is called molecular electronics.

Recently several researchers have measured charge transport in single or small groups of organic molecules connected to metal contacts [1-6]. The experiments showed that these novel molecular wire systems have a variety of interesting transport properties such as rectification,[1] negative differential resistance,[2] and switching behavior[7] to name just a few. In parallel, there have been theoretical attempts at understanding molecular conduction, both at the semiempirical[3,8-12] and first-principles[13-18] levels. Understanding molecular conduction is challenging, since it involves not just the intrinsic chemistry of the molecule, but extrinsic factors as well, such as the metal-molecule bonding geometry, contact surface microstructure and the electrostatics of the environment. Aviram and Ratner[19] were the first to propose sophisticated molecule-based systems analogous to diodes and triodes. Since their proposals, such systems have indeed been synthesized, and some techniques to connect them to external electrodes have been developed. Electronic transport through single, or at most a few, molecules adsorbed on an Au surface was observed with scanning tunneling microscope (STM), where the tip served as a counterelectrode[20,21]. STM can be used not only in the tunneling regime to image adsorbates but also in the contact regime to build few-atom nanoscopic contacts[22–25]. In addition to the STM, mechanically controllable break junctions have also provided powerful tools to study the electronic transport in metallic nanobridges or individual molecules [25–31]. A break junction is made by spanning two metallic contacts by a metallic nanowire, which is then stretched and broken using piezos. This yields two atomically sharp metallic nanocontacts. In the experiments[1,6] the break junction was exposed to a solution of the organic molecule of interest. The molecules, which were functionalized at both ends with a chemical group that binds with the metal, self-assembled onto both of the metallic tips. The solution was then evaporated, leaving behind two metal tips coated with organic molecules. The tips were then brought together, a bias voltage was applied between them and the resulting electric current that flowed from tip to tip through the self-assembled monolayer's (SAM) was measured.

Connecting a few molecules, or even a single molecule, to conducting electrodes forms the basis of a lively research activity in this field. Examples of nanowires include those constructed from aromatic dithiols,[1] carbon spheres,[32] carbon nanotubes,[33] and alkene thiols.[34] Molecular wires studied in recent years can be broadly divided into two categories; those based on *n*-alkane chains which have a large (6 eV or greater) band gap separating their highest occupied molecular orbital (HOMO) and their lowest unoccupied molecular orbital (LUMO), and those based on conjugated molecules that have a band gap of, 2–4 eV. Alkanethiol ($\text{CH}_3\text{-(CH}_2\text{)}_n\text{-SH}$) is one of the molecular systems that has been studied extensively due to its ability to form a robust SAM on gold surfaces and is useful for synthesizing insulating layers.[3, 35] Recently tunneling has been identified as the main conduction mechanism for alkanemonthiol SAMs formed in a nanometer scale junction [36]. However, the corresponding physical mechanisms governing the electron transport phenomenon in most of this experimental work are seldom mentioned. Thus we present a full atomistic *ab initio* description of the electron transport through the molecular wire especially alkanethiol system. The most common way of connecting conjugated molecules to an electron source/drain is by making a TTsubstituted molecule, *i. e.*, a TT molecule in which a hydrogen atom in its end group has been replaced by an S atom, and attaching this to an Au electrode via sulfur-gold ‘S-Au’ binding. In computational modeling of molecular nanowires, the main task has been to provide insight into how the molecule-to-metal electronic coupling and the geometry of the metal contacts affect the flow of current through the wire. In this paper, we report on first-principles calculation of the *I-V* characteristic curves of a buthane-dithiol ($\text{C}_4\text{H}_8\text{S}_2$) molecular nanowire, as shown schematically in Figure. 1. Furthermore, we have also computed the projection of the density of states (PDOS) of the combined system onto all the $\text{C}_4\text{H}_8\text{S}_2$ basis orbitals and the transmission function of the nanowire and the electrodes, and have investigated the effect of the coupling between the molecular and electrodes states.

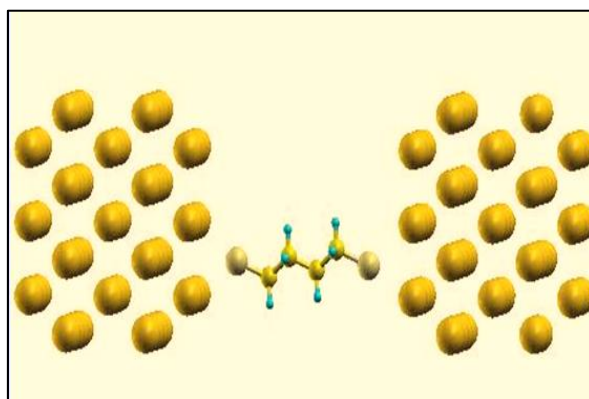


Figure.1. A conjugated molecule is chemisorbed onto the gold electrodes via the thiolate terminal groups.

The organization of the papers is as follows. We start with a brief description of the Density Functional based Non-Equilibrium Green's Function method (Section 2), and in Section 3 we present calculations of the (I-V) spectrum of a C₄H₈S₂ molecule, together with an analysis of the molecular levels important for electron transmission. In Section 4 the results are summarized.

Computational methods:

The calculations have been performed using a recently developed first-principles package SMEAGOL[37,38], which is based on the combination of DFT (as implemented in the well tested SIESTA method[39,40]) with the NEGF technique[41,42]. SMEAGOL is capable of fully selfconsistently modelling the electrical properties of nano-scale devices that consist of an atomic scale system coupling with two semi-infinite electrodes. Such nano-scale devices are referred to as two-probe systems and they are divided into three parts for theoretical calculations: left and right electrodes, and a central scattering region. The scattering region actually includes a portion of the semi-infinite electrodes. The simulation procedure of such two-probe systems can be described briefly as follows. Firstly the electronic structure of two electrodes is calculated only once by SMEAGOL to get a self-consistent potential. This potential will be shifted rigidly relative to each other by the external potential bias and provides natural real space boundary conditions for the Kohn-Sham(K-S) effective potential of the central scattering region. Then from the Green's function of the central scattering region, it can obtain the density matrix and thereby the electron density. Once the electron density is known, the DFT Hamiltonian matrix, which is used to evaluate the Green's function, can be computed using the above boundary conditions by means of standard methods.

$$\hat{G} = \lim_{\delta \rightarrow 0} \left[(E + i\delta)\hat{S} - \hat{H}_S[\rho] - \hat{\Sigma}_L - \hat{\Sigma}_R \right]^{-1} \quad (1)$$

Where \hat{H} is DFT Hamiltonian and $\hat{\Sigma}_L$ and $\hat{\Sigma}_R$ are the self-energies respectively for the left and right lead.

This procedure is iterated until self-consistency is achieved. Moreover, the current through the atomic scale system can be calculated from the corresponding Green's function and selfenergies using Landauer-Buttiker formula [43]

$$I(V) = \frac{2e}{h} \int_{-\infty}^{+\infty} dE [f_l(E - \mu_l) - f_r(E - \mu_r)] T(E, V), \quad (2)$$

where μ_L and μ_R are the electrochemical potentials of the left and right electrodes respectively,

$$\mu_L - \mu_R = eV_b$$

(3)

and f_l, f_r are the corresponding electron distribution of the two electrodes. $T(E, V)$ is the transmission coefficient at energy E and bias voltage V , which is given by

$$T(E, V) = \text{Tr}[\text{Im}\Sigma_l(E)G^R(E)\text{Im}\Sigma_r(E)G^A(E)],$$

(4)

where $G^R(E)$ and $G^A(E)$ are the retarded and advanced Green's function of the central region. Based on the eigenchannel decomposition of the conductance, this total transmission $T(E)$ can be decomposed into nonmixing eigenchannels $T_n(E)$ [44] as

$$T(E) = \sum_n T_n(E)$$

(5)

In our DFT calculation, the local-density approximation (LDA) to the exchange-correlation potential[45] is used. Only valence electrons are considered in the calculation, and the wave functions are expanded by localized numerical (pseudo)atom orbitals (PAO's)[46]. The atomic cores are described by norm-conserving pseudopotentials[47]. The structural model for our theoretical analysis is illustrated in Fig. 1. In this two-probe system, $C_4H_8S_2$ molecule couples with two atomic scale Au(100) electrodes which extend to reservoirs at $\pm\infty$ where the current is collected. Three Au atomic layers have been chosen for the electrode cell in the z-direction. In the central scattering region the $C_4H_8S_2$ molecule couples with two atomic layers to both the left and right electrode. These atomic layers in the central scattering region are large enough[48] so that the perturbation effect from the scattering region is screened and they are denoted as surface-atomic layers. In all the electrode, the gold atoms have been arranged in an ideal fcc(100) configuration with the lattice parameter of bulk gold. It has been assumed that the S's of the $C_4H_8S_2$ each placed above the hollow

sites of the electrodes. An SCF calculation by Sellers[49] has shown this to be energetically favorable for an alkanethiol. Thus the C₄H₈S₂ molecule in Fig. 1 is positioned symmetrically above the hollow sites of the two Au surfaces. Since there is no direct experimental information regarding the geometry of the molecule

and its attachment to the leads, thus we place the molecule between the electrodes assuming the Au-S distance of 2.39 Å[50].

RESULTS AND DISCUSSION

Fig. 2, shows the I-V curve of the system. We will show that the increase in the current can be understood by studying the changes of coupling between the molecular orbitals in the C₄H₈S₂ and incident states in the electrodes under various external biases.

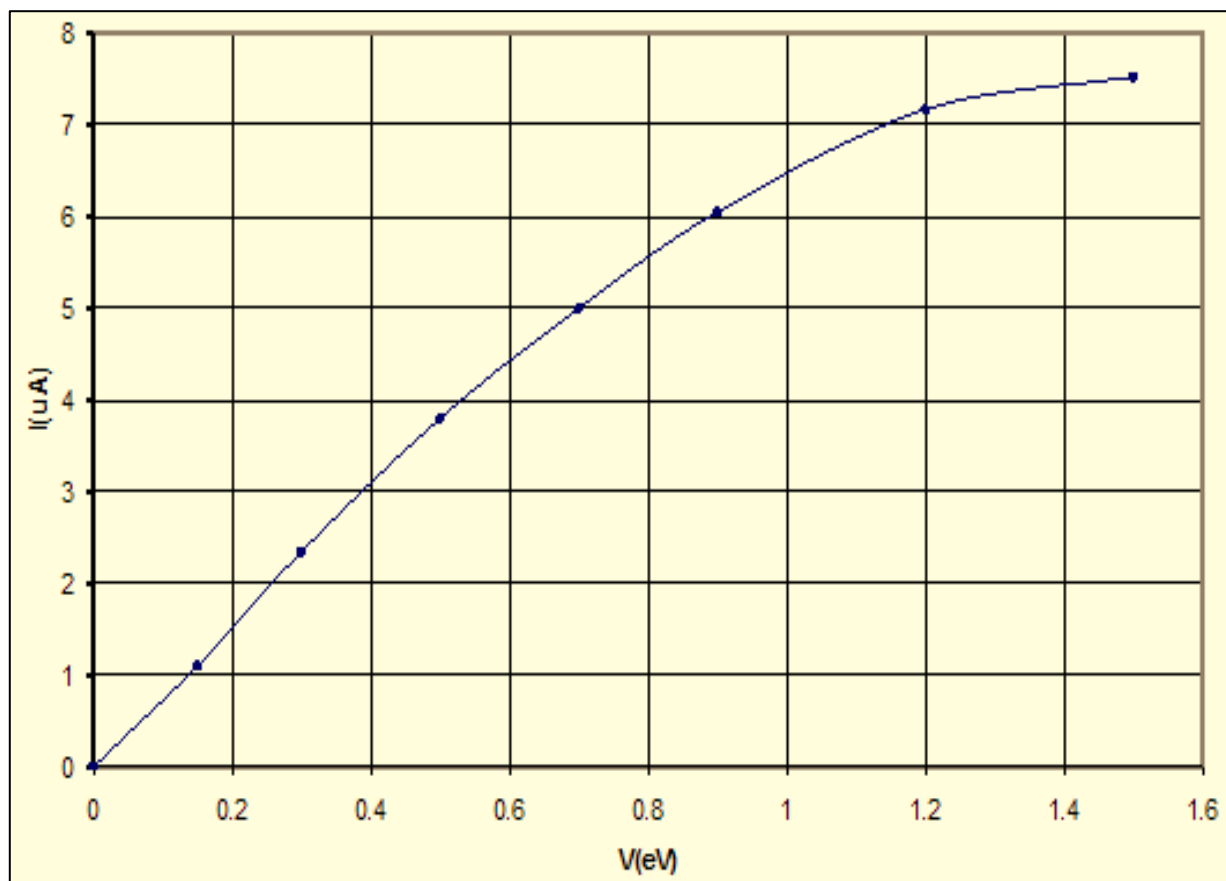


Figure. 2. Current-Voltage (I-V) curves of the two-probe system shown in figure.1.

Figure. 3 shows the zero bias transmission spectra of the system. It has a region of weak transmission above the Fermi energy, and a peak of stronger strength below, where electrons incident from one of the electrodes can transmit across the molecule to the other electrode significantly.

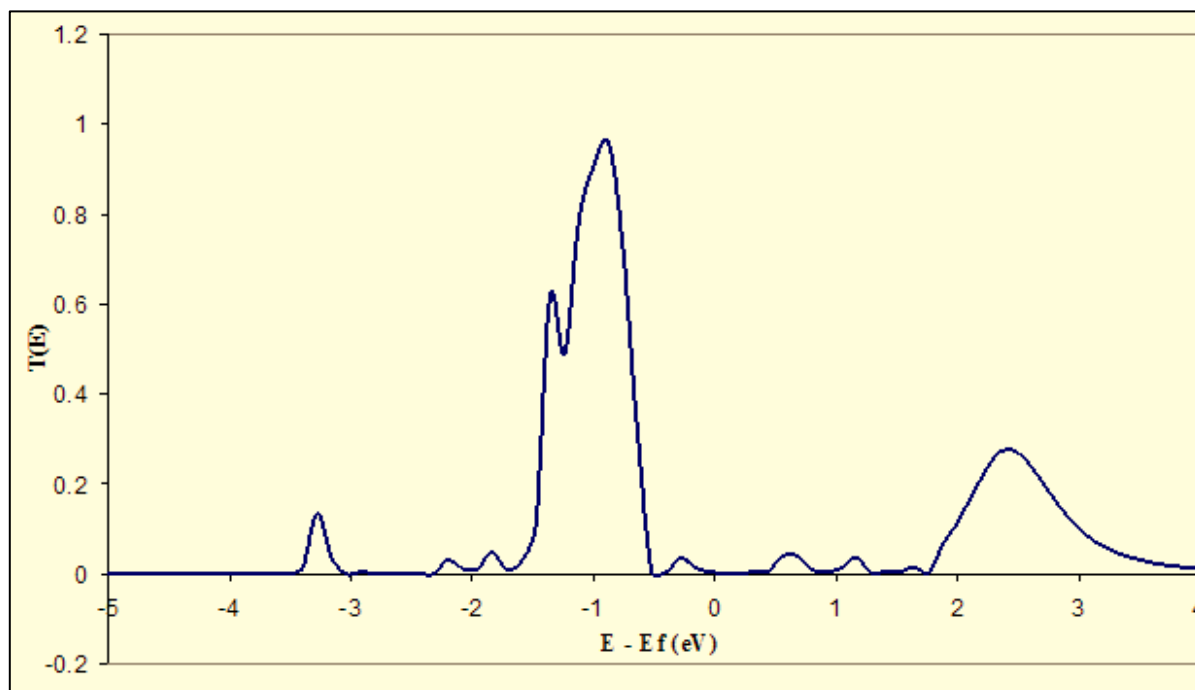


Figure. 3. a) Transmission function under zero bias as function of the injection energy of electron of the $C_4H_8S_2$ molecule coupled to Au(100) electrodes. (b) The corresponding PDOS. All energy is relative to the Fermi energy of the open system.

To understand why incident states in these two energy regions can transmit across the molecule significantly, we calculated the projection of the density of states (PDOS) of the combined system onto all the $C_4H_8S_2$ basis orbitals. The PDOS will give us information on how much the basis orbitals in the molecule contribute to the eigenstate of the whole open system and how strongly the molecule couples with the electrodes at a certain energy E . The PDOS is shown in Fig. 5.

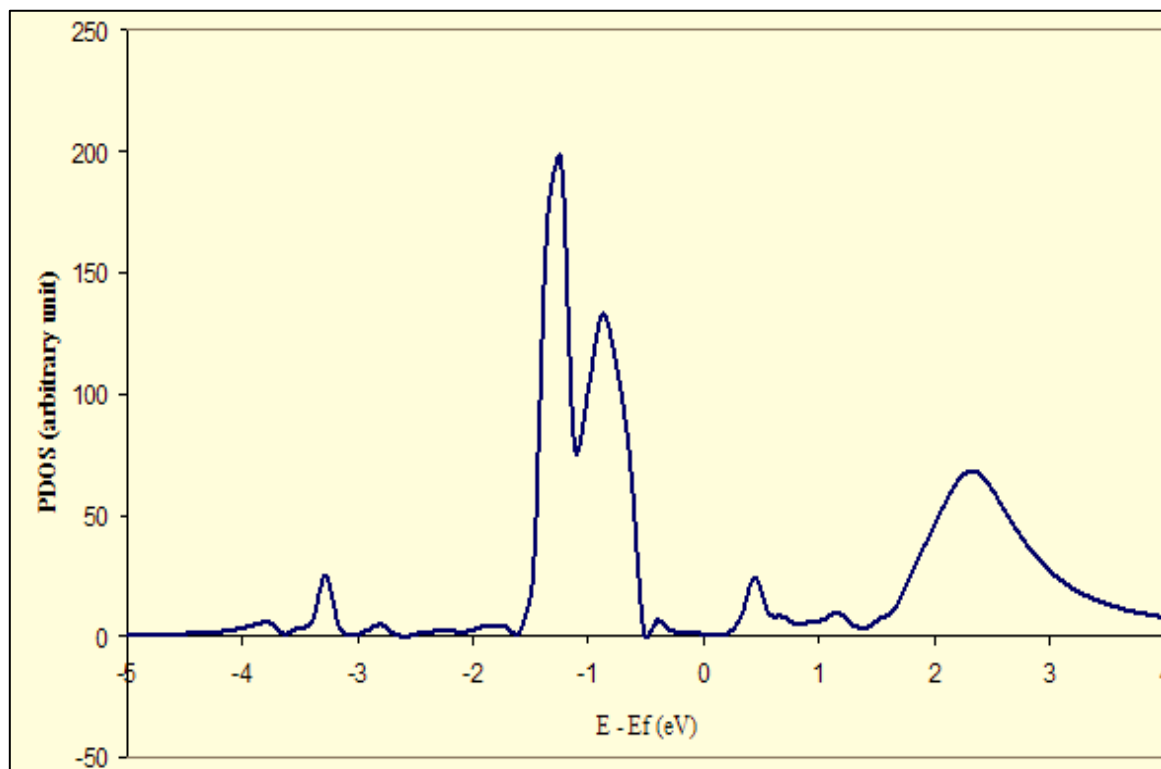


Fig. 4. The projection of the density of states (PDOS) of the combined system onto all the $C_4H_8S_2$ basis orbitals. All energy is relative to the Fermi energy of the open system.

We note that, corresponding to the T-E curve, there is also an energy region where the PDOS takes a comparatively large value. A strong coupling makes incident electrons at a certain energy easily transmit across the molecule, and this will give rise to a large transmission coefficient at this energy. This is clearly shown by comparison of the transmission spectra and PDOS spectra shown in Fig.3 and 4. As a consequence, a large transmission coefficient indicates a strong coupling between the electrodes and the molecule, and the evolution of transmission curves with external biases can help us understand how the changes of the coupling between the electrodes and molecule determines the I-V characteristics in the system. Therefore, the voltage dependence of the transmission function will be studied next.

Now we divide the whole energy region into two kinds of regions according to whether $T(E, V_b) = 0$ or not in them. In the regions where $T(E, V_b)$ is not zero the incident electrons can transmit across the molecule, and we call them transmission regions. In the other regions where $T(E, V_b)$ is zero the incident electrons cannot transmit across the molecule, and we call them transmission intervals. The current in the system is calculated by the Landauer-Buttiker formula, $I = (2e/h) \int T(E, V_b) dE$, which is transmission spectra dependent. We will now show the changes of the transmission function under various biases (Fig. 5). We see that the transmission interval near the Fermi energy under zero bias is broadened and splitted at increasing biases. It can be explained by the relative shift of the energy regions of the left and right electrodes.

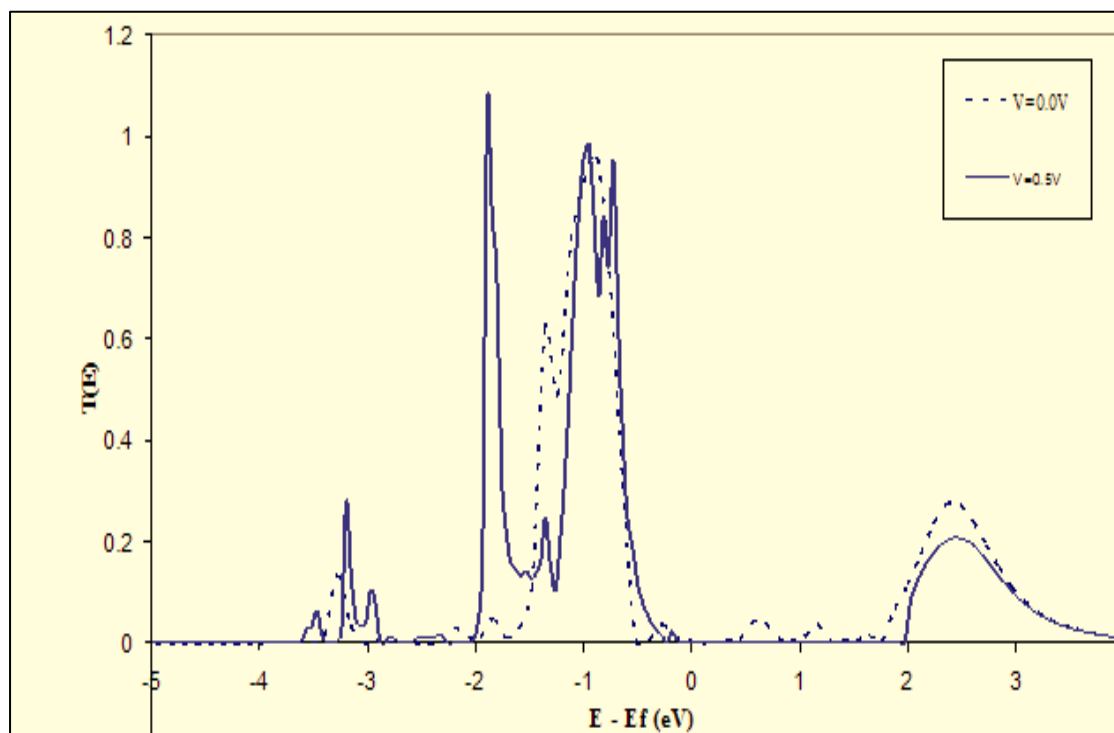


Fig. 5. Transmission coefficients under various bias voltages (solid lines indicate $V = 0.5$ V and dashed lines indicate $V = 0.0$ V) as a function of energy. Energy is relative to the Fermi energy of the open system.

Now let us see why the current was increased with increasing the bias voltage. As mentioned above the current, I , is obtained from $I(V)$ where μ_L and μ_R are the electrochemical potentials of the left/right electrodes. The region between μ_L and μ_R is called the bias window or integral window, as shown in Fig.6 with strong peaks above the the Fermi energy. Thus, the current is determined by $T(E, V_b)$ in the bias window and is further only determined by the transmission regions in the bias window because $T(E, V_b)$ is zero in the transmission interval and has no contribution to the current. It can be seen from Fig.6 that with the external bias increasing from 0.0V to, for instance, 0.5 V, the new transmission region broadens and splits rapidly since the electrons in the higher energy regions of the left electrode can transmit across the molecule to the lower energy regions of the right electrode in an increasingly wider energy region. As a result, the larger transmission coefficient indicates a strong coupling between the molecular orbitals in the $C_4H_8S_2$ and the incident states from the electrodes, thus the current increases quickly under increasing the bias voltage.

Conclusion

In conclusion, we studied the I - V characteristics in Buthane-dithiol coupled to Au(100) electrodes via the S end groups. Our study shows that the increasing in the current in this device is determined by displacement of the energy regions under various external biases. The energy regions describe the coupling of the incident states in the electrodes and the molecular orbitals in the functional molecule and are determined by features of both the molecule and the electrodes. Thus, a molecule may display different transport behaviors in different situations, for instance, with a different selection of electrodes. All these remind us that in the

design of molecular devices the information of the electrodes must also be taken into consideration very seriously.

Acknowledgment.

We thank Prof. Stefano Sanvito and Ivan Rungger for many fruitful discussions. This work was supported by the Islamic Azad University of *Yadegar Emam Khomeini Branch*.

References

- [1] Reed M.A., Zhou C., Muller C.J., Burgin T.P., and Tour J.M., *Science* 278: 252(1997).
- [2] J. Chen, M.A. Reed, A.M. Rawlett, and J.M. Tour, *Science* 286, 1550 (1999).
- [3] Tian W., Datta S., Hong S., Reifenberger R., Henderson J.I., and Kubiak C.P., *J. Chem. Phys.* 109: 2874(1998).
- [4] Collier C.P., Wong E.W., Belohradsky M., Raymo F.M., Stoddart J.F., Kuekes P.J., Williams R.S., and Heath J.R. , *Science* (1999).
- [5] Cui X.D., Primak A., Zarate X., Tomfohr J., Sankey O.F., Moore A.L., Moore T.A., Gust D., Harris G., and Lindsay S.M., *Science* 294: 571(2001).
- [6] Reichert J., Ochs R., Beckmann D., Weber H.B., Mayor M., and Lohneysen H.v. , *Phys. Rev. Lett.* 88: 176804285: 391(2002).
- [7] C.P. Collier, G. Mattersteig, E.W. Wong, Y. Luo, K. Beverly, J. Sampaio, F.M. Raymo, J.F. Stoddart, and J.R.Heath, *Science* 289, 1172 (2000).
- [8] Emberly E.G. and Kirczenow G., *Phys. Rev. B* 58: 10911(1998).
- [9] Samanta M. 1999. Master's Thesis, Purdue University.
- [10] Paulsson M. and Stafstrom S., *Phys. Rev. B* 64: 035416(2001).
- [11] Hall L.E., Reimers J.R., Hush N.S., and Silverbrook K. *J. Chem. Phys.* 112: 1510 (2000).
- [12] Cuniberti G., Fagas G., and Richter K., *Chem. Phys.* 281: 465(2002).
- [13] Damle P.S., Ghosh A.W., and Datta S. *Phys. Rev. B Rapid Comms.* 64: 201403 R(2001).
- [14] Derosa P.A. and Seminario J.M., *J. Phys. Chem. B* 105: 471(2001).
- [15] Di Ventra M., Pantelides S.T., and Lang N.D., *Phys. Rev. Lett.* 84: 979 (2000)..
- [16] Palacios J.J., P´erez-Jim´enez A.J., Louis E., and Verg´es J.A., *Phys. Rev. B* 64: 115411(2001).

- [17] Palacios J.J., P´erez-Jim´enez A.J., Louis E., SanFabi´an E., and Verg´es J.A., *Phys. Rev. B* 66: 035322 (2002).
- [18] Brandbyge M., Mozos J.-L., Ordejon P., Taylor J., and Stokbro K., *Phys. Rev. B* 65: 165401 (2002).
- [19] A. Aviram and M. A. Ratner, *Chem. Phys. Lett.* 29, 277 (1974).
- [20] S. Datta, W. Tian, S. Hong, R. Reifenberger, J. I. Henderson, and C. P. Kubiak, *Phys. Rev. Lett.* 79, 2530 (1997).
- [21] C. Joachim, J. K. Gimzewski, R. R. Schlittler, and C. Chavy, *Phys. Rev. Lett.* 74, 2102 (1995).
- [22] D. Porath, Y. Levi, M. Tarabiah, and O. Millo, *Phys. Rev. B* 56, 9829 (1997).
- [23] J. G. Hou, J. Yang¹, H. Wang, Q. Li, C. Zeng, H. Lin, W. Bing, D. M. Chen, and Q. Zhu, *Phys. Rev. Lett.* 83, 3001 (1999).
- [24] N. Agrait, J. C. Rodrigo, and S. Vieira, *Phys. Rev. B* 47, 12 345 (1993).
- [25] J. I. Pascual, J. M´endez, J. G´omez Herrero, A. M. Bar´o, N. Garc´ia, and V. T. Binh, *Phys. Rev. Lett.* 71, 1852, (1993).
- [26] C. J. Muller, J. M. Ruitenbeek, and L. J. Jong, *Phys. Rev. Lett.* 69, 140 (1992).
- [27] J. M. Krans, J. M. van Ruitenbeek, V. V. Fisun, I. K. Yanson, and L. J. de Jongh, *Nature* (London) 375, 767, (1995).
- [28] A. I. Yanson and J. M. Ruitenbeek, *Phys. Rev. Lett.* 79, 2157 (1997).
- [29] E. Scheer, N. Agrait, J. C. Cuevas, A. L. Yeyati, B. Ludoph, A. Martin-Rodero, G. R. Bollinger, J. M. van Ruitenbeek, and C. Urbina, *Nature* (London) 394, 154 (1998).
- [30] M. A. Reed, C. Zhou, C. J. Muller, T. P. Burgin, and J. M. Tour, *Science* 278, 252 (1997).
- [31] H. Park, J. Park, A. K. L. Lim, E. H. Anderson, A. P. Alivisatos, and P. L. McEuen, *Nature* (London) 407, 57, (2000).
- [32] C. Joachim, *Phys. Rev. Lett.* 74, 2102 (1995).
- [33] S. C. Tans, R. M. Verschueren, and C. Dekker, *Nature* (London) 393, 49 (1998).
- [34] R. P. Anders, J. D. Bielefeld, J. I. Henderson, D. B. Janes, V. R. Kolagunta, C. P. Kubiak, W. J. Mahoney, and R.G. Osifchin, *Science* 273, 1690 (1997).
- [35] T. Lee, W. Wang and M. A. Reed, *Jpn. J. Appl. Phys.*, 44, 523 (2005).
- [36] W. Wang, T. Lee and M. A. Reed: *Phys. Rev. B*, 68 035416(2003).
- [37] A. R. Rocha, V. M. Garc´ia-Su´arez, S. W. Bailey, C. J. Lambert, J. Ferrer, and S. Sanvito, SMEAGOL (Spin and Molecular Electronics in an Atomically-Generated Orbital Landscape. www.smeagol.tcd.ie).

- [38] A. R. Rocha, V. M. Garc'ia-Su'arez, S. W. Bailey, C. J. Lambert, J. Ferrer and S. Sanvito, *Nature Materials* 4, 335.(2005).
- [39] J. M. Soler, E. Artacho, J. D. Gale, A. Garc'ia, J. Junquera, P. Ordej'on, and D. S'anchez-Portal, J. Phys.: *Condens. Matter* 14, 2745 (2002).
- [40] SIESTA replaces core electrons by pseudopotentials and describes valence electrons by an atomic-like basis set. We use a double- ζ + polarization basis set for the organic and a single- ζ polarized basis set for the gold *s*, *p* and *d* channels. The exchange and correlation potential is calculated using the local density approximation (LDA).
- [41] A.-P. Jauho, N. S. Wingreen, and Y. Meir, *Phys. Rev. B* 50, 5528 (1994).
- [42] H. Haug, A.-P. Jauho, Quantum Kinetics in Transport and Optics of Semiconductors, Springer-Verlag, Berlin, Heidelberg, (1996).
- [43] M. BÄuttiker, Y. Imry, R. Landauer, and S. Pinhas, *Phys. Rev. B* 31, 6207 (1985).
- [44] M. Brandtge, M.R. Sørensen, and K.W Jacobsen, *Phys. Rev. B*. 56, 014956 (1997).
- [45] J.P. Perdew and A. Zunger, *Phys.Rev.B* 23, 5048 (1981).
- [46] E. Artacho, D. Suanchez-Portal, P. Ordejon, A. Garcia and J. M. Soler, *Phys. Status Solidi B* 215, 809 (1999).
- [47] N. Troullier and J.L. Martins, *Phys. Rev. B*. 43, 1993 (1991).
- [48] C. Roland, V. Meunier, B. Larade, and H. Guo, *Phys. Rev. B* 66, 035332 (2002).
- [49] H. Sellers, A. Ulman, Y. Schnidman, and J.E. Eilers, *J. Am. Chem. Soc.* 115, 9389 (1993).
- [50] Brandtge M., Mozos J.-L., Ordejon P., Taylor J., and Stokbro K. ,*Comp. Mat. Sci.* 27 ,151 (2003)